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Avogadro's Number. Gram Atomic Mass (gam) Gram Molecular Mass (gmm) The amount of a substance that contains  $6.02 \times 10^{23}$  representat.... The number of representative particles in one mole of a substa.... The atomic mass of an element expressed in grams.

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how to determine the quantitative relations between the amounts of substances involved in chemical reactions—that is, the reaction stoichiometry.

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Reactions 7.1 Multiple-Choice Questions

1) One mole of particles of any substance contains how many particles?

A) particles; B)  $3 \times$  particles; C)  $3 \times$  particles; D)  $6.02 \times$  particles; E)  $6.02 \times$  particles; Answer: D. Page Ref: 7.1

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Quantities Goals : To gain an

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understanding of : 1. Problem solving in chemistry. 2. The use of dimensional analysis to solve problems. 3. The concept of the mole. 4. The relationship between masses of substances and moles of substances. 5. The relationship between moles and the volumes and densities of ...

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### **Reactions and Quantities ...**

7.1: The Mole The mole is a key unit in chemistry. The molar mass of a substance, in grams, is numerically equal to one atom's or molecule's mass in atomic mass units. 7.1.1: Chemical Formulas as Conversion Factors; 7.1.2: New Page; 7.2: Percent Composition. 7.2.1: Determining Empirical Formulas;

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7.2.2: Determining Molecular Formulas;  
7.2.3 ...

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Chemical Reactions 7.2 -Types of  
Reactions 7.3 -Oxidation-Reduction  
Reactions 7.4 -The Mole 7.5 -Molar Mass



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and Calculations 7.6 -Mole Relationships in Chemical Equations 7.7 -Mass Calculations for Reactions 7.8 -Limiting Reactants and Percent Yield 7.9 -Energy in Chemical Reactions.

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6.7 Chapter Summary. To ensure that

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you understand the material in this chapter, you should review the meanings of the following bold terms in the following summary and ask yourself how they relate to the topics in the chapter. Chemical reactions relate quantities of reactants and products.

### **Chapter 6 - Quantities in Chemical**

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### **Reactions - Chemistry**

Question: Chapter 7 Chemical Quantities  
And Reactions Problem 7.3 - Enhanced -  
With Feedback Part A Calculate The  
Number Of C Atoms In 0.128 Mol Of C.  
Express The Number Of Atoms Of  
Carbon Numerically.  $N_C =$   
Atoms Of C Submit Request Answer Part  
B Calculate The Number Of SO,

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Molecules In 8.34 Mol Of  $\text{SO}_2$ , Express  
The Number Of Molecules Numerically.

...

### **Chapter 7 Chemical Quantities And Reactions Problem ...**

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Chemistry; Chapter 5 - Atomic Structure  
and the Periodic Table; Chapter 6 -

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Chemical Names and Formulas; Chapter 7 - Chemical Quantities; Chapter 8 - Chemical Reactions; Chapter 9 - Stoichiometry; Chemistry I Semester Exam Items; Chemistry II Semester Exam Items; Science Safety; Physics Files.

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### **Quantities**

Extending this symbolism to represent both the identities and the relative quantities of substances undergoing a chemical (or physical) change involves writing and balancing a chemical equation. Consider as an example the reaction between one methane molecule ( $\text{CH}_4$ ) and two diatomic oxygen

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molecules ( $O_2$ ) to produce one carbon dioxide ...

### **7.1: Writing and Balancing Chemical Equations - Chemistry ...**

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known as the “Twelve Principles of Green Chemistry”. One of the 12 principles is aimed specifically at maximizing the efficiency of processes for synthesizing chemical products.

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Stoichiometry. Handouts Chapters 7, 8  
and 9 Syllabus Section 7.1 and 7.2 Notes  
- Mole Conversions ... Chapter 8 #7, 10,  
22 / #17 / #20, 21 / #18, 19 Chapter 9

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Chapter 7. Chemical Reactions and Chemical Quantities ... Numerous variations on the beginning and ending computational steps are possible depending upon what particular quantities are provided and sought (volumes, solution concentrations, and so forth). Regardless of the details, all these calculations share a common

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essential component: the ...

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